



8. Two Faraday of electricity is passed through a solution of CuSO_4 . The mass of copper deposited at the cathode is: (at. mass of Cu = 63.5 amu)

- (1) 0 g (2) 63.5 g (3) 2 g (4) 127 g

Answer: We know $\text{Cu}^{++} + 2e = \text{Cu}$

Therefore 2 Faraday of electricity will deposit 1 mole of copper = 63.5 gm

Correct option is (1)